

Solubility of some ionic compounds ordered by alphabetical order

		Al ³⁺	NH ₄ ⁺	Ba ²⁺	Ca ²⁺	Cu ²⁺	H ⁺	Fe ²⁺	Fe ³⁺	Pb ²⁺	Li ⁺	Mg ²⁺	Hg ₂ ²⁺	Hg ²⁺	Ni ²⁺	K ⁺	Ag ⁺	Na ⁺	Zn ²⁺
Acetate	CH ₃ COO ⁻	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S
Bromide	Br ⁻	S	S	S	S	S	S	S	S	X	S	S	X	S	S	S	X	S	S
Carbonate	CO ₃ ²⁻	X	S	X	X	X	-	X	X	X	S	X	X	X	X	S	X	S	X
Chlorate	ClO ₃ ⁻	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S
Chloride	Cl ⁻	S	S	S	S	S	S	S	S	X	S	S	X	S	S	S	X	S	S
Fluoride	F ⁻	X	S	X	X	S	S	X	X	X	S	X	-	-	S	S	S	S	S
Hydroxide	OH ⁻	X	S	S	S	X	-	X	X	X	S	X	X	X	X	S	X	S	X
Iodide	I ⁻	S	S	S	S	-	S	S	S	X	S	S	X	S	S	S	X	S	S
Nitrate	NO ₃ ⁻	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S
Phosphate	PO ₄ ³⁻	X	S	X	X	X	S	X	X	X	X	X	X	X	X	S	X	S	X
perchlorate	ClO ₄ ⁻	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S
Sulfate	SO ₄ ²⁻	S	S	X	X	S	S	S	S	X	S	S	X	-	S	S	S	S	S
Sulfide	S ²⁻	-	S	-	X	X	S	X	-	X	S	-	X	X	X	S	X	S	X
Sulfite	SO ₃ ²⁻	X	S	X	X	X	S	X	X	X	S	S	-	-	X	S	X	S	X
Oxide	O ²⁻	X	-	S	S	X	S	X	X	X	-	X	X	X	X	S	X	-	X

Solubility of some ionic compounds ordered by solubilities

		NH ₄ ⁺	Li ⁺	Na ⁺	K ⁺	H ⁺	Ba ²⁺	Ca ²⁺	Pb ²⁺	Hg ₂ ²⁺	Ag ⁺	Hg ²⁺	Cu ²⁺	Ni ²⁺	Zn ²⁺	Mg ²⁺	Fe ²⁺	Fe ³⁺	Al ³⁺
Acetate	CH ₃ COO ⁻	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S
Nitrate	NO ₃ ⁻	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S
Chlorate	ClO ₃ ⁻	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S
perchlorate	ClO ₄ ⁻	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S	S
Chloride	Cl ⁻	S	S	S	S	S	S	S	X	X	X	S	S	S	S	S	S	S	S
Bromide	Br ⁻	S	S	S	S	S	S	S	X	X	X	S	S	S	S	S	S	S	S
Iodide	I ⁻	S	S	S	S	S	S	S	X	X	X	S	-	S	S	S	S	S	S
Sulfate	SO ₄ ²⁻	S	S	S	S	S	X	X	X	X	S	-	S	S	S	S	S	S	S
Sulfite	SO ₃ ²⁻	S	S	S	S	S	X	X	X	-	X	-	X	X	X	S	X	X	X
Sulfide	S ²⁻	S	S	S	S	S	-	X	X	X	X	X	X	X	X	-	X	-	-
Hydroxide	OH ⁻	S	S	S	S	-	S	S	X	X	X	X	X	X	X	X	X	X	X
Oxide	O ²⁻	-	-	-	S	S	S	S	X	X	X	X	X	X	X	X	X	X	X
Carbonate	CO ₃ ²⁻	S	S	S	S	-	X	X	X	X	X	X	X	X	X	X	X	X	X
Phosphate	PO ₄ ³⁻	S	X	S	S	S	X	X	X	X	X	X	X	X	X	X	X	X	X

Fluoride is removed

S = soluble X = not soluble - = unstable, decomposes or does not exist

note: For this table, slightly soluble compound are considered as insoluble